

Organic Chemistry
Cyclic Aliphatic Compounds
Practice Set

- 1) (13.1) Draw structural formulas of:
- a) Methylcyclopentane
 - b) 1-methylcyclohexene
 - c) 3-methylcyclopentene
 - d) *trans*-1,3-dichlorocyclobutane
 - e) *cis*-1-bromo-2methylcyclopentane
 - f. cyclohexylcyclohexane
 - g. cyclopentylacetylene
 - h. 4-chloro-1,1-dimethylcycloheptane
 - i. bicyclo[2.2.1]hepta-2,5-diene
 - j. 1-chlorobicyclo[2.2.2]octane
- 2) (13.2) Give structures and names of the principal organic products expected from each of the following reactions:
- a. cyclopropane + Cl₂, FeCl₃
 - b. cyclopropane + Cl₃ (300.[°]C)
 - c. cyclopropane + conc. H₂SO₄
 - d. cyclopentane + Cl₂, FeCl₃
 - e. cyclopentane + Cl₂ (300. [°]C)
 - f. cyclopentane + conc. H₂SO₄
 - g. cyclopentene + Br₂/CCl₄
 - h. cyclopentene + Br₂ (300. [°]C)
 - i. 1-methylcyclohexene + HCl
 - j. 1- methylcyclohexene + Br₂ (aq)
 - k. 1-methylcyclohexene + HBr (peroxides)
 - l. 1,3-cyclohexadiene + HCl
 - m. cyclopentanol + H₂SO₄ (heat)
 - n. bromocyclohexane + KOH(alc)
 - z. 1-methylcyclopentene + cold conc. H₂SO₄
 - aa. 3-methylcyclopentene + O₃, then H₂O/Zn
 - bb. 1-methylcyclohexene + (BH₃)₂; H₂O, OH⁻
 - cc. 1-methylcyclohexene + Hg(OAc)₂, H₂O;NaBH₄
 - o. cyclopentene + cold KMnO₄
 - p. cyclopentene + HCO₂OH
 - r. cyclopentene + hot KMnO₄
 - s. cyclopentene + NBS
 - t. 3-bromocyclopentene + KOH (hot)
 - u. 1,4-cyclohexanediol + H₂SO₄ → C₁₂H₂₀
 - w. cyclopentene + CHCl₃ + *t*-BuOK
 - x. cyclopentene + CH₂I₂ +Zn(Cu)
 - y. chlorocyclopentane + (C₂H₅)₂CuLi
- 3) Outline all steps in the laboratory synthesis of each of the following from cyclohexanol.
- a) cyclohexene
 - b) cyclohexane
 - c) *trans*-1,2-dibromocyclohexane
 - d) *cis*-1,2 cyclohexanediol
 - e) *trans*-1,2-cyclohexanediol
 - f) OHC(CH₂)₄CHO
 - g. adipic acid, HOOC(CH₂)₄COOH
 - h. bromocyclohexane
 - i. 2-chlorocyclohexanol
 - j. 3-bromocyclohexene
 - k. 1,3cyclohexadiene
 - l. cyclohexylcyclohexane
 - m. *norcarane*, bicyclo[4.1.0]heptane